

Name _____

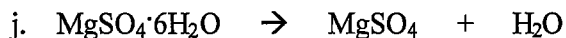
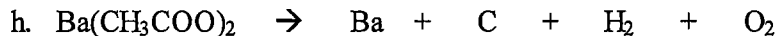
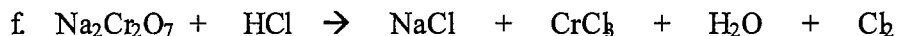
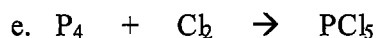
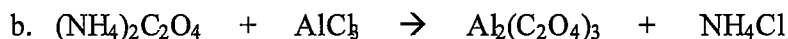
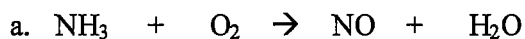
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Due Date _____

Chemistry 11**Assignment – Chemical Equations**

This Assignment will be marked and you are allowed to do one set of corrections.

1. Balance the following equations (1 mark each = 10 marks)



2. Write a balanced chemical equation for each of the following. Don't forget *diatomic* elements! (2 marks each = 24 marks)

a. aluminum metal reacts with bromine to form aluminum bromide

Answer _____
|

- b. hydrochloric acid neutralizes aluminum hydroxide to form water & aluminum chloride

Answer _____

- c. hexane (C_6H_{14}) burns in oxygen to produce carbon dioxide and water

Answer _____

- d. carbon dioxide and water are reacted to produce glucose ($C_6H_{12}O_6$) and oxygen in photosynthesis

Answer _____

- e. aluminum nitrate reacts with lithium sulphate to form aluminum sulphate and lithium nitrate

Answer _____

- f. ammonia and phosphoric acid react to form ammonium phosphate

Answer _____

- g. nitrogen dioxide reacts with water to form nitric acid and nitrogen monoxide

Answer _____

- h. bromine reacts with sodium iodide to produce iodine and sodium bromide

Answer _____

- i. calcium reacts in water to produce hydrogen gas and calcium hydroxide

Answer _____

- j. butanol (C_4H_9OH) burns in oxygen to produce carbon dioxide and water

Answer _____

- k. hydrogen peroxide decomposes to form water and oxygen gas

Answer _____

- l. aluminum bromide reacts with ammonium dichromate to produce aluminum dichromate and ammonium bromide

Answer _____

Review

Name _____

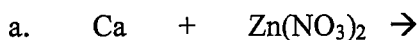
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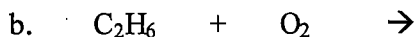
Chemistry 11**Assignment – Completing, Balancing & Classifying
Chemical Equations**

This Assignment will be marked and you are allowed to do one set of corrections.

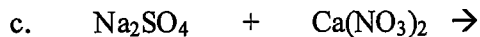
1. **Complete, balance and classify** the following equations as *synthesis, decomposition, single replacement, double replacement, neutralization or combustion*.
(3 marks each = 51 marks)



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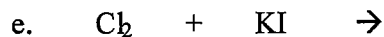
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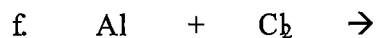
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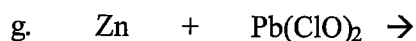
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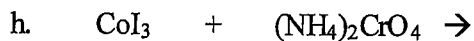
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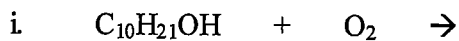
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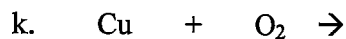
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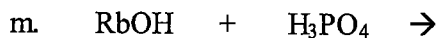


(Assume combining capacity of Cu is 2+)

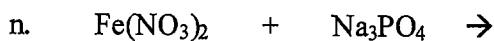
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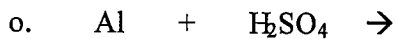
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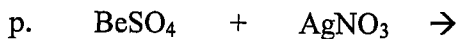
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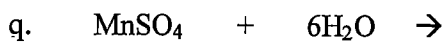
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