

Write the formulas for the following ionic compounds:

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| 21) | lithium acetate | $\text{LiC}_2\text{H}_3\text{O}_2$ |
| 22) | iron (II) phosphate | $\text{Fe}_3(\text{PO}_4)_2$ |
| 23) | titanium (II) selenide | TiSe |
| 24) | calcium bromide | CaBr_2 |
| 25) | gallium chloride | GaCl_3 |
| 26) | sodium hydride | NaH |
| 27) | beryllium hydroxide | $\text{Be}(\text{OH})_2$ |
| 28) | zinc carbonate | ZnCO_3 |
| 29) | manganese (VII) arsenide | Mn_3As_7 |
| 30) | copper (II) chlorate | $\text{Cu}(\text{ClO}_3)_2$ |
| 31) | cobalt (III) chromate | $\text{Co}_2(\text{CrO}_4)_3$ |
| 32) | ammonium oxide | $(\text{NH}_4)_2\text{O}$ |
| 33) | potassium hydroxide | KOH |
| 34) | lead (IV) sulfate | $\text{Pb}(\text{SO}_4)_2$ |
| 35) | silver cyanide | AgCN |
| 36) | vanadium (V) nitride | V_3N_5 |
| 37) | strontium acetate | $\text{Sr}(\text{C}_2\text{H}_3\text{O}_2)_2$ |
| 38) | molybdenum sulfate | $\text{Mo}(\text{SO}_4)_3$ |
| 39) | platinum (II) sulfide | PtS |
| 40) | ammonium sulfate | $(\text{NH}_4)_2\text{SO}_4$ |